

**BHU M.Sc. CHEMISTRY ENTRANCE - 2011**

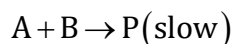
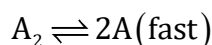
- The change in entropy when two moles of a monatomic perfect gas is compressed to half its volume and simultaneously heated to twice its initial temperature is  
(a)  $R \ln 2$                       (b)  $3 R \ln 2$                       (c)  $5 R \ln 2$                       (d)  $7 R \ln 2$
- A heat engine operates between 100 K and 600 K. The heat discharged into the cold sink in a reversible process when 5 kJ of heat is supplied by the hot source, is  
(a) 2 kJ                      (b) 2.5 kJ                      (c) 3 kJ                      (d) 5.5 kJ
- For which of the following processes  $q = 0$ ,  $w = 0$ ,  $\Delta U = 0$  and  $\Delta H = 0$ ?  
(a) Reversible isothermal process in a perfect gas  
(b) Reversible adiabatic process in a perfect gas  
(c) Adiabatic expansion of a perfect gas into vacuum  
(d) Reversible constant-volume process in a perfect gas
- The fugacity of a certain gas at 200 K and 50 bars is 25 bars. The difference of its chemical potential from that of a perfect gas in the same state is  
(a)  $200 R \ln 25$                       (b)  $200 R \ln 2$                       (c)  $-200 R \ln 2$                       (d)  $-200 R \ln 25$
- The entropy of mixing 1 mole of hexane with 1 mole of heptane at 298 K is  $11.4 \text{ JK}^{-1}$ . The Gibbs energy of mixing (assuming the solution to be ideal) is  
(a) -1.72 kJ                      (b) -3.44 kJ                      (c) 1.72 kJ                      (d) 3.44 kJ
- The number of degrees of freedom of the system  $\text{KClO}_3(\text{s}) \rightleftharpoons 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  is  
(a) zero                      (b) one                      (c) two                      (d) three
- The expression that relates the partial molar properties of the components in a mixture is known as  
(a) van't Hoff equation                      (b) Gibbs-Duhem equation  
(c) Duhem-Margules equation                      (d) Raoult's law
- For a very dilute electrolyte solution with  $\gamma_{\pm} < 1$ ,  $\gamma_{\pm}$  should increase with increase in  
(a) solvent's density                      (b) solvent's dielectric constant  
(c) ionic strength                      (d) cationic charge



9. Which of the following statements is not correct?
- (a)  $\Delta G_{\text{mix}}$  at constant temperature and pressure must be negative
  - (b) Intermolecular interactions are negligible in an ideal solution
  - (c) Solute-solute interactions are negligible in an ideally dilute solution
  - (d) Activity coefficients are never negative
10. Which of the following statements is correct?
- (a) If  $\Delta G^\circ > 0$ , no amount of products can be obtained when the reaction is run at constant temperature and pressure
  - (b) It is possible for the entropy of a closed system to decrease substantially in an irreversible process
  - (c) In any closed system with P-V work only, G is always minimized at equilibrium
  - (d)  $\Delta(TS) = T\Delta S + S\Delta T$
11. The condition for the attainment of phase-equilibrium in a closed electrochemical system is the equality of
- (a) surface potentials (b) electric potentials
  - (c) chemical potentials (d) electrochemical potentials
12. The direction of a chemical reaction at constant temperature and pressure is the direction of
- (a) decrease of Gibbs free energy of the system
  - (b) decrease of Helmholtz free energy of the system
  - (c) increase of entropy of the system
  - (d) decrease of enthalpy of the system
13. The quantum yield of the photochemical decomposition of HI
- $$\text{HI} + h\nu \rightarrow \text{H} + \text{I}$$
- $$\text{H} + \text{HI} \rightarrow \text{H}_2 + \text{I}$$
- $$\text{I} + \text{I} + \text{M} \rightarrow \text{I}_2 + \text{M}$$
- with respect to HI is
- (a) 0.5 (b) 1 (c) 1.5 (d) 2



14. HI is absorbed strongly on gold. Assuming Langmuir isotherm to apply, the order of the reaction  $2\text{HI} \rightarrow \text{H}_2 + \text{I}_2$  on gold is
- (a) zero                      (b) 0.25                      (c) 0.5                      (d) 1
15. Which of the following statements is not correct for Langmuir isotherm?
- (a) It applies to monolayer adsorption
- (b) Under conditions  $\theta \ll 1$  and  $\alpha$  (Freundlich exponent)  $\approx 1$ , it reduces to Freundlich isotherm
- (c) It applies to dissociative
- (d) It applies to chemisorption
16. The activation energy and entropy of a bimolecular gas phase reaction at 600 K are  $200 \text{ kJ mol}^{-1}$  and  $-200 \text{ JK}^{-1} \text{ mol}^{-1}$  respectively. The free energy of activation is
- (a)  $70 \text{ kJ mol}^{-1}$               (b)  $80 \text{ kJ mol}^{-1}$               (c)  $310 \text{ kJ mol}^{-1}$               (d)  $320 \text{ kJ mol}^{-1}$
17. Among the following statements which is the correct one?
- (a) The heat of chemisorption is always larger than that of physisorption
- (b) Langmuir isotherm specifically assumes the existence of active centres
- (c) Promoters are themselves catalysts
- (d) Increase in surface tension with concentration leads to negative adsorption
18. The activation energy of the gas-phase association between  $\text{F}_2$  and  $\text{IF}_5$ , a first-order reaction in each of the reactants, is  $58.6 \text{ kJ mol}^{-1}$ . The activation enthalpy at 340 K is
- (a)  $53 \text{ kJ mol}^{-1}$               (b)  $55.8 \text{ kJ mol}^{-1}$               (c)  $58.6 \text{ kJ mol}^{-1}$               (d)  $61.4 \text{ kJ mol}^{-1}$
19. In photochemical reaction  $\text{A} \rightarrow 2\text{B} + \text{C}$ , the quantum efficiency with 500 nm light is  $2 \times 10^2 \text{ mol Einstein}^{-1}$ . After exposure of 300 m moles of A to the light, 2 m moles of B is formed. The number of photons absorbed by A is
- (a)  $1 \times 10^{18}$               (b)  $3 \times 10^{18}$               (c)  $6 \times 10^{18}$               (d)  $9 \times 10^{18}$
20. The condition for which the reaction rate of an enzymolysis that follows Michaelis-Menten kinetics, is half its maximum value, is
- (a)  $[\text{S}] \ll K_M$               (b)  $[\text{S}] = K_M$               (c)  $[\text{S}] = K_M / 2$               (d)  $[\text{S}] \gg K_M$
21. For the mechanism



the reaction order with respect to  $A_2$  is

- (a) 0                      (b) 0.5                      (c) 1                      (d) 2

22. Which of the following relations does not hold for the activity ( $A$ ) of a radioactive substance?

(a)  $\frac{A}{A_0} = \left(\frac{1}{2}\right)^{t/t_{1/2}}$

(b)  $\frac{A}{A_0} = \exp\left(-0.693 \frac{t}{t_{0.5}}\right)$

(c)  $t_{0.5} / t_{0.1} = \ln 2$

(d)  $\frac{A}{A_0} = 1 - 0.693 \frac{t}{t_{0.5}}$  at  $t \rightarrow 0$

23. A powder diffraction photograph from tungsten shows lines which indices as (110), (200), (211), (220), (310), (222), (321), (400),.... The symmetry of the unit cell is

- (a) primitive              (b) end-centred              (c) face-centred              (d) body-centred

24. Among the following halides which one forms van der Waals crystals?

- (a) NaCl                      (b)  $\text{BeCl}_2$                       (c)  $\text{HgCl}_2$                       (d) HF

25. A form of  $\text{CaCO}_3$  (c) has orthorhombic lattice with  $a = 5.0 \text{ \AA}$ ,  $b = 8.0 \text{ \AA}$ ,  $c = 5.6 \text{ \AA}$  and density =  $3.0 \text{ gm cm}^{-3}$  at room temperature. The number of  $\text{Ca}^{2+}$  ions per unit cell of the crystal is

- (a) 1                      (b) 2                      (c) 4                      (d) 8

26. The ratio of the translational partition functions of  $\text{D}_2$  and  $\text{H}_2$  at the same temperature and volume is

- (a) 2                      (b) 1.414                      (c) 2.83                      (d) 4

27. In which of the following systems is the energy levels separation the largest?

- (a) an electron in a radical in a field of 0.300 T  
(b) a  $^{14}\text{N}$  nucleus in 600 MHz NMR spectrometer  
(c) a proton in the same spectrometer



- (d) a deuteron in the same spectrometer
28. Which of the following functions is not an eigen function of  $\frac{d^2}{dx^2}$ ?
- (a)  $\cos kx$                       (b)  $\exp(-kx^2)$                       (c)  $kx$                       (d)  $\exp(ikx)$
29. The ratio of mean molar masses of a given polymer sample as determined by light scattering, sedimentation and osmotic pressure measurement methods is
- (a) 1:1:2                      (b) 2:1:2                      (c) 1:2:1                      (d) 2:2:1
30. In milk at 37°C *Lactobacillus acidophilus* has a generation time of 75 minutes. What is the population relative to the initial value at 150 minutes?
- (a) 4.0                      (b) 2.25                      (c) 2.0                      (d) 1.75
31. If the pressure of a gas at constant temperature is doubled, the viscosity of the gas will be
- (a) quadrupled                      (b) doubled                      (c) halved                      (d) unchanged
32. Among the following molecules which one shows pure rotation spectra?
- (a)  $N_2$                       (b)  $H_2S$                       (c)  $CO_2$                       (d)  $CH_4$
33. The SI unit of radiation dose is
- (a) becquerel                      (b) curie                      (c) rad                      (d) gray
34. It is found that a particle in a one-dimensional box of length of  $L$  can be excited to  $n = 2$  state from the ground state by the light of frequency  $\nu$ . If the box length is doubled, the frequency needed to produce the above transition becomes
- (a)  $\nu/4$                       (b)  $\nu/2$                       (c)  $2\nu$                       (d)  $4\nu$
35. For a hydrogen atom in an  $n = 4$  state, the maximum possible z-component of orbital angular momentum is
- (a)  $2\hbar$                       (b)  $3\hbar$                       (c)  $\sqrt{12}\hbar$                       (d)  $\sqrt{6}\hbar$
36. Which one of the following is an acceptable approximate wave function for a state of the helium atom?
- (a)  $[1s(1)1s(2) - 1s(1)1s(2)][\alpha(1)\beta(2)]$
- (b)  $[1s(1)1s(2)][\alpha(1)\alpha(2) + \beta(1)\beta(2)]$



- (c)  $[1s(1)2s(2) + 2s(1)1s(2)][\alpha(1)\alpha(2)]$   
(d)  $[1s(1)2s(2) + 2s(1)1s(2)][\alpha(1)\beta(2) - \beta(1)\alpha(2)]$

37. Which one of the following statements concerning  $H_2^+$  is incorrect?
- (a) The non-degenerate LCAO-MOs (without spin) must be either symmetric or anti-symmetric for inversion  
(b) The lowest energy MO (without spin) of the molecule is anti-symmetric for inversion  
(c) The ground state has a multiplicity of two  
(d) The MOs transform into AOs of the helium ion as the two nuclei are fused together
38. Which one of the following is the correct formula for the lowest energy eigen function for a particle in a one-dimensional box having infinite barriers at  $x = -L/2$  and  $L/2$ ?
- (a)  $\sqrt{\frac{2}{L}} \sin\left(\frac{\pi x}{L}\right)$     (b)  $\sqrt{\frac{2}{L}} \cos\left(\frac{\pi x}{L}\right)$     (c)  $\sqrt{\frac{2}{L}} \exp\left(\frac{i\pi x}{L}\right)$     (d)  $\sqrt{\frac{2}{L}} \exp\left(\frac{-i\pi x}{L}\right)$
39. Which of the following equations is used to calculate the number of theoretical plates?
- (a)  $\frac{t_R - t_0}{t_0}$     (b)  $16\left(\frac{t_R}{W}\right)^2$   
(c)  $\frac{k_2}{k_1}$     (d)  $2(t_{R_2} - t_{R_1}) / (W_1 + W_2)$
40. Which of the following techniques is based on selectively inducing radioactivity and measuring the emitted radiation?
- (a) Isotope dilution analysis    (b) Radiometric titration  
(c) Neutron activation analysis    (d) All of the above
41. Which of the following techniques can be used only for volatile compounds?
- (a) Gas chromatography    (b) HPLC  
(c) Ion chromatography    (d) All of the above
42. Which of the following refers to ion exchange capacity
- (a) Nature of exchanging ions



- (b) Nature of strong cation exchanger  
(c) Nature of strong anion exchanger  
(d) Total number of ion active groups per unit length of material
43. In electrogravimetry of cations the working electrode is  
(a) anode (b) cathode  
(c) both cathode and anode (d) neither anode nor cathode
44. Which of the following techniques is/are feasible approach in the determination of a substance that cannot be isolated in pure form for gravimetry or for determination by other methods?  
(a) Neutron activation analysis (b) Isotope dilution analysis  
(c) Radiometric titration (d) All of the above
45. Which of the following frequencies corresponds to carbonyl stretch vibration in acids?  
(a)  $1625\text{ cm}^{-1}$  (b)  $1715\text{ cm}^{-1}$  (c)  $1745\text{ cm}^{-1}$  (d)  $1800\text{ cm}^{-1}$
46. A particular vibration in a polyatomic molecule is IR active if during vibration, there is a change in  
(a) polarizability (b) dipole moment (c) frequency (d) potential energy
47. Moisture in a drug can be determined by  
(a) Malaprade reagent (b) EDTA reagent  
(c) Karl Fischer reagent (d) chloramine-T reagent
48. Which one is more toxic?  
(a) Hg (b)  $(\text{CH}_3)_2\text{Hg}$  (c)  $\text{Hg}^{2+}$  (d)  $\text{Hg}_2^{2+}$
49. The most efficient technique for the separation of amino acids is  
(a) adsorption chromatography (b) partition chromatography  
(c) ion-exchange chromatography (d) paper chromatography
50. Which one is not a pollutant?  
(a) CO (b)  $\text{CO}_2$  (c)  $\text{SO}_3$  (d)  $\text{NO}_2$
51. Which one is the sink of  $\text{CO}_2$ ?  
(a) Plant (b) Ocean (c) Air (d) Soil



52. Ozone layer is a protective shield against  
(a) visible light      (b) ultraviolet light      (c) infrared rays      (d) cosmic rays
53. Nessler's reagent is  
(a)  $\text{KHgI}_4$       (b)  $\text{K}_2\text{HgI}_4 + \text{NH}_4\text{OH}$   
(c)  $\text{K}_2\text{HgI}_4$       (d)  $\text{KHgI}_4 + \text{NH}_4\text{OH}$
54. Which of the following statements is not correct?  
(a) In instrumental methods of analysis visual indicators are not required  
(b) In classical volumetric titrations visual indications are not required  
(c) In classical qualitative analysis removal of interfering radicals is essential  
(d) Spectroanalytical techniques can be applied for qualitative and quantitative analysis
55. Visible region in the electromagnetic spectrum extends from  
(a) 200-400 nm      (b) 800-900 nm      (c) 400-4000  $\text{cm}^{-1}$       (d) 400-800 nm
56. Hard water can be softened in  
(a) a calorimeter      (b) a chromatograph  
(c) an ion exchanger      (d) an earthen ware pot
57. The technique used for the separation of components of a mixture is called  
(a) chromatography      (b) IR spectroscopy  
(c) electronic spectroscopy      (d) polarography
58. Essential constituent of an amalgam is  
(a) Ag      (b) Hg      (c) Mg      (d) Fe
59. Potential of which of the following electrodes does not depend on pH of the solution?  
(a) Glass electrode      (b) Hydrogen electrode  
(c) Quinhydrone electrode      (d) calomel electrode
60. The unit of equivalent conductance is  
(a)  $\text{ohm}^{-1} \text{cm}^{-1}$       (b)  $\text{ohm}^{-1} \text{cm}^{-2}$   
(c)  $\text{ohm}^{-1} \text{cm}^2$       (d) None of the above
61. Which quantity remains unchanged on changing temperature?





- (a) Mole fraction (b) Molarity  
(c) Normality (d) None of the above
62. The strongest ligand in spectro-chemical series is  
(a)  $F^-$  (b)  $OH^-$  (c)  $H_2O$  (d)  $CO$
63. The iron is rusted, then it is  
(a) oxidized (b) reduced (c) evaporated (d) decomposed
64. Two elements cannot be combined chemically to make  
(a) a compound (b) another element  
(c) a gas (d) a liquid
65. Choose the law that states, "effective chemical changes are brought about by absorbed radiations only"  
(a) Beer's law (b) Grotthus-Draper law  
(c) Lambert's law (d) law of photochemical equivalence
66. What quantity of limestone on heating will give 56 kg of  $CaO$ ?  
(a) 10 kg (b) 55 kg (c) 44 kg (d) 100 kg
67. The unit of absorptivity is  
(a)  $cm^{-1} g^{-1}L$  (b)  $cm^{-1} mol^{-1}L$  (c)  $cm^{-1} g L^{-1}$  (d) No unit
68. Oxine is a  
(a) precipitating reagent (b) redox reagent  
(c) drug (d) dye
69. Free radicals may be detected by  
(a) mass spectrometry (b) NMR spectroscopy  
(c) ESR spectroscopy (d) infrared spectroscopy
70. Lead can be the best analyzed colorimetrically using which of the following reagents?  
(a) Dithizone (b) EDTA (c) DMG (d) Acetyl acetone
71. Which of the following is a hydride ion donor?  
(a) NAD (b) ATP (c) NADH (d) Coenzyme A
72. Na and K can be estimated using



- (a) X-ray fluorescence (b) IR spectroscopy  
(c) ion selective electrode (d) All of the above
73. How many mL of concentrated HCl (density = 1.18 g/mL, 36.0 wt % HCl, MW = 36.5) should be diluted to 1.00 L to produce a 0.100 M solution?  
(a) 8.59 mL (b) 85.9 mL (c) 0.859 mL (d) 17.18 mL
74. A  $1.00 \times 10^{-3}$  M solution of HCl includes  ${}^{36}_{17}\text{C}$  ( $t_{1/2} = 3.09 \times 10^5$  y). The specific activity (the activity per unit mass or unit volume) of the solution is 5000 decays  $\text{min}^{-1}$   $\text{mL}^{-1}$ . What fraction of Cl atoms in solution are  ${}^{36}\text{Cl}$  atoms?  
(a) 0.194 (b) 0.0194 (c) 1.940 (d) 0.00194
75. A  $6.44 \times 10^{-4}$  M solution of Ti-peroxide complex had  $T = 0.340$  when measured in a 1.00 cm cell at 410 nm. What is the concentration of Ti-peroxide complex in a solution that had  $T = 0.516$  when measured at 410 nm in a 0.500 cm cell?  
(a)  $7.89 \times 10^{-2}$  M (b)  $15.78 \times 10^{-2}$  M (c)  $7.89 \times 10^{-4}$  M (d)  $3.95 \times 10^{-3}$  M
76. The partition coefficient for ethyl iodide, E, between octanol and water is  
$$K_p = \frac{[E]_o}{[E]_w} \approx 100$$
  
What percent of E present in 50.00 mL of water would remain if the water was extracted with 10.00 mL of octanol?  
(a) 47.6% (b) 4.76% (c) 9.52 (d) 95.2%
77. Arrange the ionization energies of the following ions in increasing order  
(A)  $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+}$   
(B)  $\text{Mn}^{2+} \rightarrow \text{Mn}^{3+}$   
(C)  $\text{Fe}^{+1} \rightarrow \text{Fe}^{2+}$   
Choices are  
(a)  $A < B < C$  (b)  $C < B < A$  (c)  $C < A < B$  (d)  $B < C < A$   
(At. No. of Fe and Mn are 26 and 25)



78. In the first row transition metal ion ( $2+$ ) having four electrons in 3d orbitals (with 4s and 4p orbitals empty), the metal ion forms a hypothetical diamagnetic tetrahedral complex  $[MA_4]^{2-}$  with  $A^-$  anion. The electronic configuration of the  $M^{2+}$  ion in the complex is
- (a)  $[Ar]3d_{xy}^2 3d_{xz}^2 3d_{yz}^0 3d_{z^2}^0 3d_{x^2-y^2}^0 4s^0 4p^0$       (b)  $[Ar]3d_{xy}^0 3d_{xz}^0 3d_{yz}^2 3d_{z^2}^2 3d_{x^2-y^2}^0 4s^0 4p^0$
- (c)  $[Ar]3d_{xy}^0 3d_{xz}^0 3d_{yz}^0 3d_{z^2}^2 3d_{x^2-y^2}^2 4s^0 4p^0$       (d)  $[Ar]3d_{xy}^1 3d_{xz}^1 3d_{yz}^2 3d_{z^2}^0 3d_{x^2-y^2}^0 4s^0 4p^0$
79. Out of the following which has the least tendency to form  $M = O$  species?
- (I)  $Sc = O$       (II)  $P = O$       (III)  $Ln = O$       (IV)  $Ac = O$
- (a)  $Sc = O$  and  $P = O$       (b)  $Ln = O$
- (c)  $Ac = O$       (d)  $Sc = O, P = O$  and  $Ac = O$
- ( $Ln =$  Lanthanides and  $Ac =$  actinides)
80. The magnetic properties of lanthanides are largely determined by the formula
- (a)  $\mu_{eff} = \sqrt{n(n+2)}$
- (b)  $\mu_{eff} = \sqrt{n(n+2)} + \text{some contribution of orbital contribution}$
- (c)  $\mu_{eff} = \sqrt{L(L+2)}$
- (d)  $\mu_{eff} = g_J \sqrt{J(J+1)}$ , where  $g_J = \frac{3}{2} + \frac{S(S+1) - L(L+1)}{2J(J+1)}$
81.  $Ln^{2+}$  ions are largely coloured while  $Ln^{3+}$  ions are not coloured. The colour of  $Ln^{2+}$  compounds are due to transition
- (a)  $4f \rightarrow 5d$
- (b)  $4f \rightarrow 4f$
- (c) charge transfer from ligand to metal ion
- (d)  $5d \rightarrow 5d$  transition



82. Transition metals (some ions) form complexes with molecular  $N_2$ . There are two modes of coordination of  $N_2$  in the complexes of the type  $[M(N_2)(L)_n]$ . These modes are (a) end-on and (b) side-on. The stability of these complexes in these two modes is
- (a) end-on more stable than side-on  
(b) side-on more stable than end-on  
(c) both are highly unstable (the relative stability depends on ligands)  
(d) both are equally stable  
(L = ligand)
83. The structure of hyponitrous acid molecule ( $H_2N_2O_2$ ) is
- (a) linear (A–B–C–D type)                      (b) bent ( $\sphericalangle$  type)  
(c) square planar ( $\square$  type)                      (d) triangular ( $\triangle$  type)
84. We have two complexes  $[Mn(H_2O)_6]^{3+}$  ion ( $d^4$  system) and  $[Re(H_2O)_6]^{3+}$  ion ( $d^4$  system). Their CFSE and the pairing energy (per pair of electrons) are given below
- $[Mn(H_2O)_6]^{3+} - 10Dq = 250 \text{ kJ/mol}$ ; Pairing energy = 300 kJ/mol  
 $[Re(H_2O)_6]^{3+} - 10Dq = 400 \text{ kJ/mol}$ ; Pairing energy = 180 kJ/mol
- Predict from their data, the complexes are
- (a) both high spin                                      (b) Mn-high spin and Re-low spin  
(c) both low spin                                      (d) Mn-low spin, Re-high spin
85.  $[NiCl_2(PPh_3)_2]$  is paramagnetic with  $\mu_{\text{eff}} = 2.9 \text{ BM}$ . All the four ligands are monodentate. The geometry of the molecule is
- (a) square planar                                      (b) octahedral  
(c) square pyramidal (d) tetrahedral



86. There is no d-d transition in  $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$  ( $d^5$  system). This is high spin complex. The absence of any transition in the visible region is due to breaking of
- (a) Laporte
  - (b) Parity selection rule ( $g \rightarrow g$  and  $u \rightarrow u$  not allowed)
  - (c) spin selection rule (spin of the electron cannot be changed during transition)
  - (d) All the above three rules (1, 2 and 3)
87. According to the IUPAC system which is the correct formula for the complex [diaquadibromodi(methylamine)] cobalt (III) nitrate?
- (a)  $[\text{Co}(\text{H}_2\text{O})_2\text{Br}_2(\text{MeNH}_2)_2]\text{NO}_3$
  - (b)  $[\text{CoBr}_2(\text{H}_2\text{O})_2(\text{MeNH}_2)_2]\text{NO}_3$
  - (c)  $[\text{Co}(\text{H}_2\text{O})_2(\text{MeNH}_2)_2\text{Br}_2]\text{NO}_3$
  - (d)  $[\text{Co}(\text{MeNH}_2)_3(\text{H}_2\text{O})_2\text{Br}_2]\text{NO}_3$
88.  $[\text{Cr}(\text{H}_2\text{O})_3]^{3+}$  is violet whereas  $[\text{Cr}(\text{NH}_3)_6]^{3+}$  is yellow. The wavelength absorbed by  $[\text{Cr}(\text{NH}_3)_6]^{3+}$  as compared to that by  $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$  in their absorption spectra will be
- (a) higher
  - (b) lower
  - (c) same
  - (d) These complexes will not show any absorption in the visible region
89. Although fluorine is better oxidizing agent than oxygen, but  $\text{Mn}_2\text{O}_7$  exists and  $\text{MnF}_7$  does not. This is because of
- (a) steric crowding and difficult to fitting seven fluoride ions around Mn centre
  - (b) very high electron affinity of fluorine as compared to oxygen
  - (c) very low lattice energy of  $\text{MnF}_7$  compared to oxygen compound
  - (d) very low bond energy of Mn–F bond compared to Mn–O bond
90. Covalent radius of gold (125 pm) is less than that of silver (1.33 pm). This is because of
- (a) transition metal contraction

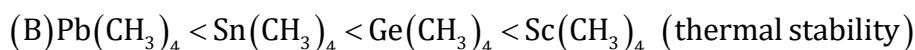
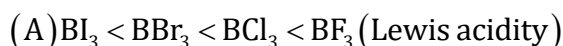


- (b) lanthanide contraction  
(c) lack of shielding of 4d orbitals compared to 5d orbitals  
(d) relatively less effective shielding by 3 p and 3s orbitals in silver as compared to that by 4s, 4p and 4d
91. Using VSEPR model, the shape of the following molecules  $\text{IF}_5^{-2}$  (A),  $\text{XeF}_4$  (B),  $\text{IO}_3^-$  (C),  $\text{BrF}_3$  (D) are
- (a) (A) pentagon shape, (B) square planar, (C) trigonal pyramid, (D) T-shape  
(b) (A) trigonal bipyramid, (B) tetrahedral, (C) planar, (D) planar  
(c) (A) square pyramid, (B) tetrahedral, (C) trigonal pyramid, (D) planar  
(d) (A) pentagonal pyramid, (B) tetrahedral, (C) trigonal pyramid (D) planar
92. Arrange the hydrides of 15-group elements ( $\text{NH}_3, \text{PH}_3, \text{AsH}_3, \text{SbH}_3 < \text{BiH}_3$ ) in increasing contribution of p-orbital of X in the X-H bond. Choose the correct alternative from the following alternatives given below
- (a)  $\text{Bi} < \text{Sb} < \text{As} < \text{P} < \text{N}$  (b)  $\text{P} < \text{N} < \text{As} < \text{Bi} < \text{Sb}$   
(c)  $\text{N} < \text{P} < \text{As} < \text{Sb} < \text{Bi}$  (d)  $\text{Bi} < \text{Sb} < \text{N} < \text{P} < \text{As}$
93. What are A, B and C in the following reactions?
- (I)  $\text{B}_2\text{H}_6 + \text{NH}_3 \rightarrow (\text{A})$   
(II)  $\text{B}_2\text{H}_6 + \text{H}_2\text{O} \rightarrow (\text{B})$   
(III)  $\text{B}_2\text{H}_6 \xrightarrow[\text{in the absence of air}]{\text{Heated at } 100^\circ\text{C}} (\text{C})$
- Here A, B and C respectively are
- (a) borazine,  $\text{B}_2\text{O}_3$  and  $\text{B}_4\text{H}_8$  (b)  $2\text{BH}_3 \cdot \text{NH}_3, \text{H}_3\text{BO}_3 + \text{H}_2, \text{B}_{10}\text{H}_{14}$   
(c)  $\text{BN}, \text{H}_3\text{BO}_3$  only,  $\text{B}_4\text{H}_8$  (d)  $\text{BN}, \text{H}_3\text{BO}_3 + \text{H}_2, \text{B}_4\text{H}_8$
94. Which one of the following contain  $(3c-2e)$  bonds?



(a) A, B and C (b) A and C (c) A, C and D (d) B and C

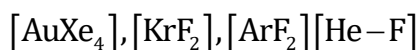
95. Which one of the following orders of two properties of 14- and 15-group elements is not correct?



The correct alternative from the following ones is

(a) B (b) A  
(c) No one is wrong (d) Both are wrong

96. Which ones of the following compounds do not exist?



Choose the correct choice out of the following

(a)  $[\text{AuXe}_4]$  and  $[\text{He}-\text{F}]$  (b)  $[\text{KrF}_2]$  and  $[\text{ArF}_2]$   
(c)  $[\text{Ar}-\text{F}_2]$  and  $[\text{He}-\text{F}]$  (d)  $[\text{ArF}_2]$ ,  $[\text{He}-\text{F}]$ ,  $[\text{AuXe}_4]$

97. Alkalides and electrides are crystalline compounds of

(a) alkali metals in (+1) oxidation states and electrons respectively  
(b) alkali metals in (-1) oxidation states and electron acting as anion  
(c) methyl group and electron acting as anion respectively  
(d) methyl group and alkali metal group

98. Among the molecules,  $\text{BiF}_3$ ,  $\text{BiCl}_3$ ,  $\text{BiBr}_3$  and  $\text{BiI}_3$ , the one which is most coloured is

(a)  $\text{BiF}_3$  (b)  $\text{BiBr}_3$  (c)  $\text{BiCl}_3$  (d)  $\text{BiI}_3$

99.  $[\text{CrO}_8]^{3-}$  ion is known. This ion probably is

(a) complex of peroxide ion with  $\text{Cr}^{5+}$  ion  
(b) polymeric complex with oxide ion acting as bridges



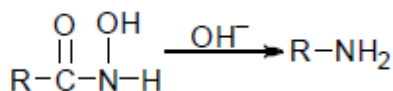




105. In borazine ( $B_3N_3H_6$ ) molecule, the number of isomers which are possible of its disubstituted borazine molecule of the formula  $[B_3N_3H_4X_2]$  without changing its ring structure is
- (a) one                      (b) two                      (c) four                      (d) six
106. Solution of alkali metals in liquid  $NH_3$  conducts electricity. It is due to formation of
- (a)  $Na^+ + Na^-$  ion in liquid ammonia  
(b)  $Na^+ + e^-(NH_3)_x$  in liquid ammonia  
(c)  $Na^-$ ,  $(NH_2)^-$  and  $NH_4^+$  ion in liquid  $NH_3$   
(d) The solution conducts like a metallic conductor with solvated electron carrying the charge
107. A compound alloy of metals P and Q has a unit cell containing P atom at the corners, while atom Q are present at the face centres. The formula of the compound should be
- (a) PQ                      (b)  $PQ_2$                       (c)  $PQ_3$                       (d)  $P_3Q$
108. A complex of a certain metal ion has a magnetic moment of 4.90 BM (4 unpaired electrons) and another complex of the same metal ion in the same oxidation state is diamagnetic. The complex is octahedral. The central metal could be one of the following
- ( $Cr^{3+}$ ,  $Mn^{2+}$ ,  $Fe^{2+}$ ,  $Mn^{3+}$ ,  $Fe^{3+}$ ,  $Co^{2+}$ )
- $Cr^{3+}$  ion is  $[Ar]4d^3$  system. The metal ion is
- (a)  $Mn^{2+}$                       (b)  $Fe^{2+}$                       (c)  $Co^{2+}$                       (d)  $Fe^{3+}$
109. The enthalpy of hydration ( $\Delta H$ ) of  $Cr^{2+}$  [ $d^4$  system octahedral complex ( $Cr(H_2O)_6$ ) $^{2+}$ ] is  $-460$  kcal/mol. In the absence of CFSE, this value of  $\Delta H$  is ( $-435$ ) kcal/mole. The value of  $10 Dq$  (or  $\Delta_0$ ) is
- (a)  $14600\text{ cm}^{-1}$                       (b)  $21500\text{ cm}^{-1}$                       (c)  $9525\text{ cm}^{-1}$                       (d)  $25252\text{ cm}^{-1}$
- ( $1\text{ kcal} = 350\text{ cm}^{-1}$ )



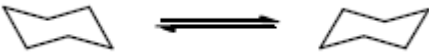
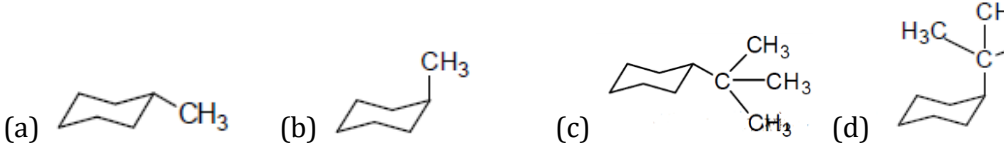

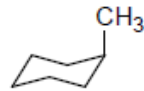
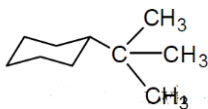
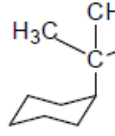
110. Which of the  $d^n$  ion will have the smallest CFSE if  $\Delta$  (or  $10Dq$ ) is greater than pairing energy  $d^6, d^7, d^8, d^9, d^{10}$ ?
- (a)  $d^6, d^{10}$                       (b)  $d^7, d^9$                       (c)  $d^{10}$                       (d)  $d^5, d^8, d^{10}$
111. During excessive physical stress (like running), the body makes A which reacts with water to form its conjugate base and B ions. These, in turn, lowers the pH in the muscles which causes pain in the muscles. Here A and B respectively are
- (a) citric acid and  $H_3O^+$                       (b) lactic acid and  $H_3O^+$
- (c) enzyme and  $NH_4^+$                       (d) gluconic acid and  $NH_4^+$
112. Which ones of the following are closed packed structures?
1. AABB AABB...                      2. ABAC ABAC...
3. AB AB AB...                      4. ABC ABC ABC
- (a) 1, 2, 3                      (b) 1, 3, 4                      (c) 2, 3, 4                      (d) 2, 3
113. Out of following actinides (I) Np, (II) Pu, (III) Es, (IV) Am one having +2 oxidation state as the only relatively important state is
- (a) Np                      (b) Pu                      (c) Es                      (d) Am
114. Which is not an ambident nucleophile?
- (a)  $CN^-$                       (b)  $SCN^-$                       (c)  $NO_2^-$                       (d) DMSO
115. What is the increasing order of stability of following carbocations (give least stable first)?
- (I) Tropylium cation    (II)  $CH_2 = CH - CH_2^+$
- (III)  $(C_6H_5)_2 C^+$                       (IV)  $CH_3^+$
- (a) III < I < II < IV    (b) IV < II < III < I    (c) I < III < II < IV    (d) IV < III < II < I
116. The type of rearrangement

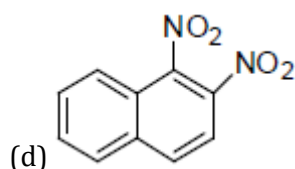
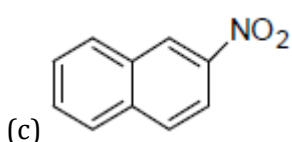
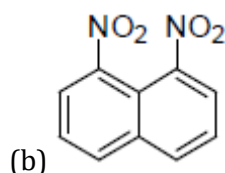
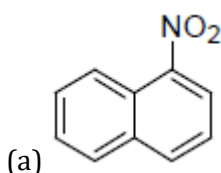
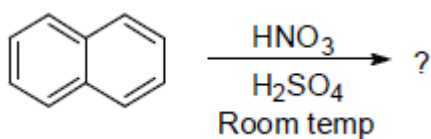




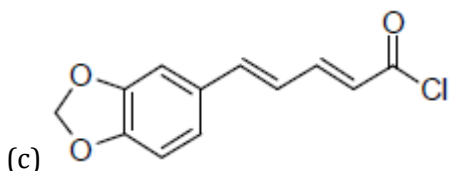
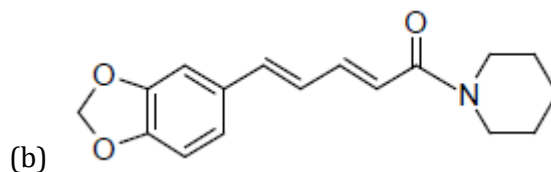
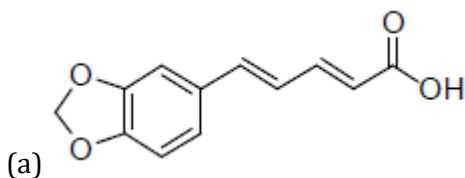




- (b) boat < twist boat < half chair < chair  
(c) chair < half chair < twist boat < boat  
(d) chair < twist boat < boat < half chair
130. In cyclohexane molecule, when one chair conformation changes to other
- 
- (a) all hydrogens originally axial become equatorial  
(b) potential energy of chair conformation increases  
(c) angle strain reduces  
(d) None of the above
131. Which of the following is least stable?
- 
- (a)  (b)  (c)  (d) 
132. Which of the following is not an alkaloid?  
(a) Quinine (b) Reserpine (c) Camphor (d) Piperine
133. Indigo dye is deep blue colored due to  
(a) conjugation of double bond and non-planarity of molecule  
(b) conjugation of double bond and planarity of molecule  
(c) hydrogen bonding and planarity of molecule  
(d) None of the above
134. A powerful anti-malarial agent obtained from cinchona bark is  
(a) emetine (b) piperine (c) quinine (d) colchicine
135. A polynuclear compound having three benzene rings fused angularly is  
(a) naphthalene (b) anthracene (c) phenanthrene (d) chrysene
136. IUPAC name of nicotine is  
(a) naphthalene (b) anthracene (c) phenanthrene (d) chrysene
137. Predict the product for the following reaction



138. Structure of piperine is



(d) None of these

139. In Chichibabin reaction, base used is

(a)  $\text{KNH}_2$

(b)  $\text{NaNH}_2$

(c)  $\text{C}_6\text{H}_5\text{NH}_2$

(d)  $\text{NH}_3$

140. Sulphonation in indole takes place at

(a) position-2

(b) position-3

(c) both

(d) None of these

141. Isoquinoline on oxidation with alkaline  $\text{KMnO}_4$  gives

(a) phthalic acid

(b) benzoic acid

(c) cinchomeric acid (d) both 1 and 3

142.  $\alpha$ -Terpeneiol gives p-Cymene on treatment with

(a)  $\text{HNO}_3$

(b)  $\text{H}_2\text{SO}_4$

(c)  $\text{HCl}$

(d)  $\text{CH}_3\text{COOH}$



143. The sweetest sugar amongst the following is  
(a) lactose                      (b) fructose                      (c) glucose                      (d) sucrose
144. The bond that determines the secondary structure of protein is  
(a) coordinate bond                      (b) covalent bond  
(c) hydrogen bond                      (d) ionic bond
145. The main structural feature of protein is  
(a) ester linkage                      (b) ether linkage                      (c) peptide linkage                      (d) ionic linkage
146. Buna-S is obtained by the polymerization of butadiene and  
(a) chloroprene                      (b) styrene                      (c) acrylonitrile                      (d) adipic acid
147. Terylene is a polymer of  
(a) adipic acid and hexamethylene diamine  
(b) terephthalic acid and ethylene glycol  
(c) phenol and formaldehyde  
(d) adipic acid and ethylene glycol
148. Which is an example of thermosetting polymer?  
(a) Polyethene                      (b) Neoprene  
(c) Polyvinyl chloride (d) Bakelite
149. Which of the following is chain growth polymer?  
(a) Glyptal                      (b) Nylon-66                      (c) Nylon-6                      (d) Polypropylene
150. Alkaloids are found in  
(a) seeds                      (b) root                      (c) bark                      (d) All of the above